



Versatile Applications of Mannich Base Ligands and their Metal Complexes: A Review Article

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Abstract

Mannich base is a versatile compound that can be easily modified to introduce different functional groups, allowing for the creation diverse selection of items with varying features. Additionally, the Mannich reaction is a valuable tool in organic synthesis, due to the fact it provides an effortless and efficient approach for synthesizing C-N bonds. Overall, The Mannich base and even its derivatives are essential in many aspects of chemistry and its complexes are in the pharmaceutical industry. Studies have revealed that it shows good anti-cancer, anti-mycobacterial, remarkable anti-HIV, anti-tubercular, anti-convulsant, anti-fungal, antiviral, antitumor, cytotoxic activities and in industrial applications such as in the creation of polymers, surface activity agents, detergents and resins. The presence of the basic Mannich sidechain has shown marked antimalarial, anti-inflammatory, analgesic and antimicrobial activities. These compounds have also been shown to inhibit corrosion, as well as antioxidant and reducing agents. This review article shows the definition, importance and different applications of Mannich base ligands with transitional metal. These complexes exhibit potent anti-microbial, antiviral, and anti-cancer activities, showcasing their potential in pharmaceutical research and drug development. Moreover, the luminescent properties of Mannich base metal complexes have been harnessed for applications in optoelectronics and sensing. Their tunable emission profiles make them suitable candidates for various sensing platforms and light-emitting devices Mannich base metal complexes.

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1. Introduction

The Mannich reaction, discovered in 1917 [1-6] stands out as a prominent multi-component reaction, involving the simultaneous addition of aldehyde, amine, and ketone in a one-pot condensation to produce amino carbonyl compounds, commonly known as Mannich bases. This reaction holds significant importance as it facilitates the creation of new C-C bonds, enabling the synthesis of diverse amino carbonyl molecules, see Figure 1.

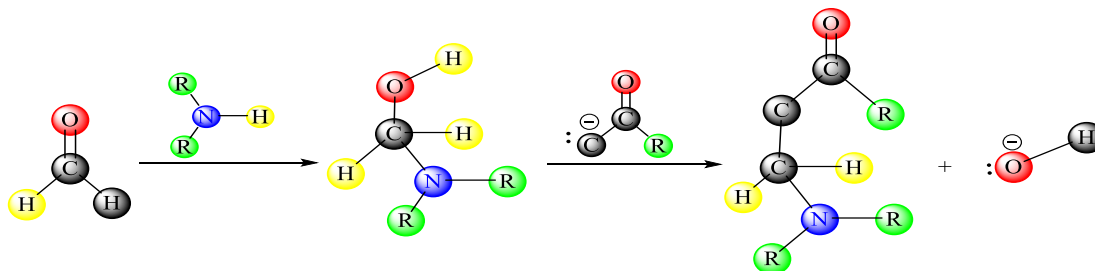


Figure 1: general structure of Mannich bases.

Consequently, the Mannich reaction has a significant effect in the outcome synthesis of essential biomolecules such as peptides, amino alcohols, lactam, amino acids, and other biologically significant compounds. Its versatility and efficiency in forming key chemical bonds make it a fundamental and valuable method in modern organic synthesis, the compounds obtained could be versatile intermediates that help in this [7-12]. In recent years, there has been extensive research on the metallic complexes of Mannich bases, primarily due to the ligands' remarkable selectivity towards different transition metal ions. Notably, these complexes have demonstrated a strong inclination to form, leading to compounds with significant anti-cancer and anti-malarial properties. The straightforward synthesis process, adaptability, and diverse applications have sustained Mannich base complexes as a prominent and widely studied area of research. Their role in advancing coordination chemistry is evident from the existing literature, highlighting the vital contribution of these metallic complexes to the field's development [13-18]. Extensive volume on the toxicological, biological, and chemical features of Mannich bases with broad applications as polymers, dispersants in lubricants, and pharmaceutical agents. This is demonstrated by a review of the literature on Mannich reactions. Studying the structural and binding features of different Mannich base complexes can play a significant role in better understanding complex biological processes. [19-22]. A few of the Mannich bases have been studied as corrosion inhibitors in our group. Through the nitrogen atoms π e⁻ and the aromatic ring contained in the Mannich bases molecules, inhibition can be induced by their readily absorbed onto the metal surface. Therefore, as mild steel corrosion preventive in an acidic media, Mannich bases which have been successfully examined, can be used [23-27]. Mannich bases have found diverse and practical applications in various fields, such as treating conventional macro-molecular materials like textiles, leather, paper and synthesizing polymers. They are widely used as additives in petroleum production, cosmetics, water treatment, dyes, and analytical reagents. However, medicinal chemistry is the most significant and noteworthy application of the Mannich reaction. Initially, Mannich bases have shown intriguing biological activities, and with diligent screening processes, many more of their potential medicinal properties are yet to be uncovered. The continuous publication of numerous research papers each year supports the growing interest in exploring the medicinal aspects of Mannich bases and their potential as valuable candidates for drug development and therapeutic interventions. Second, amino methylation of drugs can be used to improve the delivery of drugs to the human body [28-31].

2. Literature Review of Synthesis Mannich Base Ligands and Their Complexes

2.1. [1-(di-n-butylamine-N-methyl)-mercapto-1H-benzimidazole (L₁) and 1-(diphenylamine-N-methyl)-mercapto-1H-benzi-midazole] (L₂)

Synthesis of Mannich base ligands, L₁ and L₂ and reacted with metal ions to produce complexes (Co, Ni, Cu, and Zn). To examine the improvement of the reaction, the thin-layer method technique purified thin-layer chromatography utilized. Various spectroscopic techniques were used to describe the Mannich base ligands and their metal complexes, including ¹H/NMR, FT/IR, UV/Visible, ICPES, ¹³C/NMR and AAS. NMR and FT/IR analyses confirmed the monoanionic bi-dentate co-ordination mode of L₁ and L₂, while the tetrahedral geometries of the metal complexes were proposed based on AAS/ICP, electronic spectroscopic and magnetic moment results. Additionally, the luminescent behaviour of the synthesized compounds was investigated, revealing broad emission bands that indicate the involvement of charge transfer transitions [32], see in Figure 2.

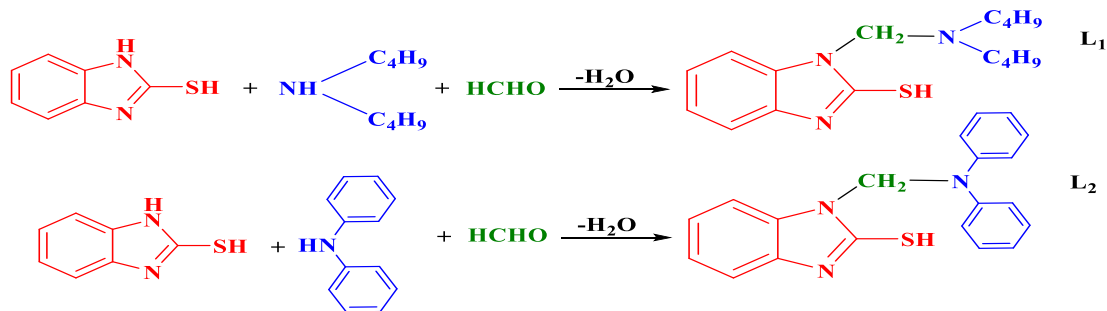


Figure 2: synthesis of two Mannich base ligands [32].

2.2. [1-((diallylamino)methyl)-1*H*-benzo[d]imidazol-2-ol (A1) and 1-(((4-chlorophenyl)amino)methyl)-1*H*-benzo[d]imidazol-2-ol (A2)]

Mannich base compounds (A1) and (A2) were formed and described by reaction *O*-phenylenediamine and urea with various amine. The hydrogen atom on the nitrogen atom was then swapped through the Mannich reaction employing primary and secondary amines. The synthetic chemicals were then examined using H/NMR and FT/IR spectroscopy for their structural organization [33], see Figure 3.

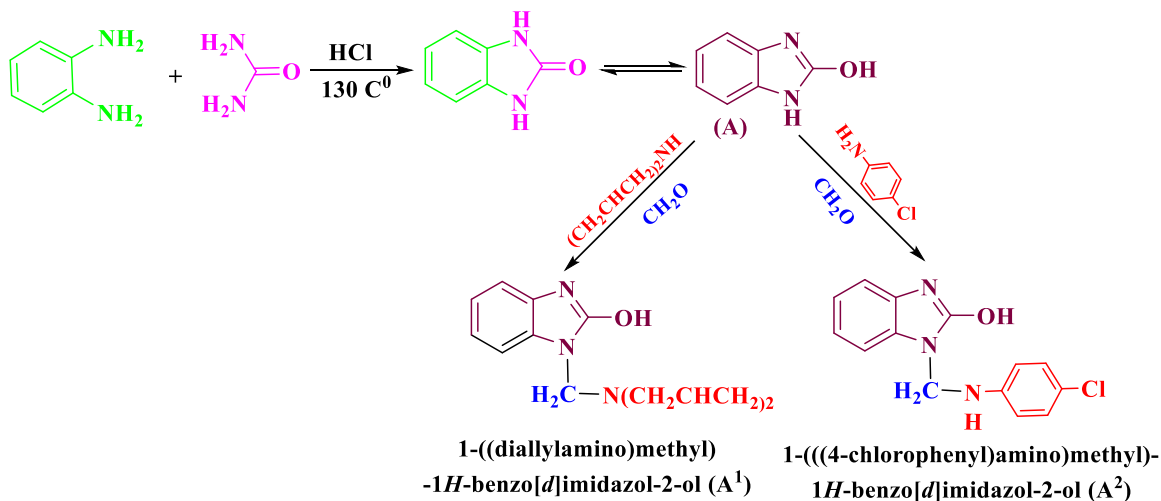


Figure 3: synthesis of compounds [33].

2.3. [Co₃(L₂)(NO₃)₄].CH₃CN (1) and [Ni₃(L₂)(NO₃)₄].CH₃CN (2)

Produced trinuclear cobalt (1) and nickel (2) compounds. These complexes were made using the ligand H₂L₂, which was made by reacting HL₁ as Mannich base in a [tow: one] mole ratio with propylene diamine. Reacting 5-bromo-salicylaldehyde with N, N, N'-trimethyl-ethylene diamine affords the Mannich base HL₁ precursor. In the ensuing complexes, metal ions are connected by NO and OC₆H₅ atoms. Two of each compound's 4 nitrate groups behave as chelating ligands towards the terminal metal ions. Compounds 1 and 2 have their core metal ions accommodated within the salen-like cavity of the organic ligand. End-on azido or end-on nitrate bridges can be applied in place of nitrate bridges; N-thiocyanato groups are formed, resulting in the compounds [Ni₃(L₂)(N₃)₂(NO₃)₂] (3), [Co₃(L₂)(NCS)₂(NO₃)₂] (5) and [Co₃(L₂)(N₃)₂(NO₃)₂] (4). The magnetic characteristics of these five compounds were examined, and it was discovered that compounds 1 and 2 had intramolecular antiferromagnetic interactions, whereas compounds 3, 4, and 5 exhibit ferromagnetic connections. The nature and amplitude of these interactions are addressed and compared to prior research on similar systems [34] see Figure 4.

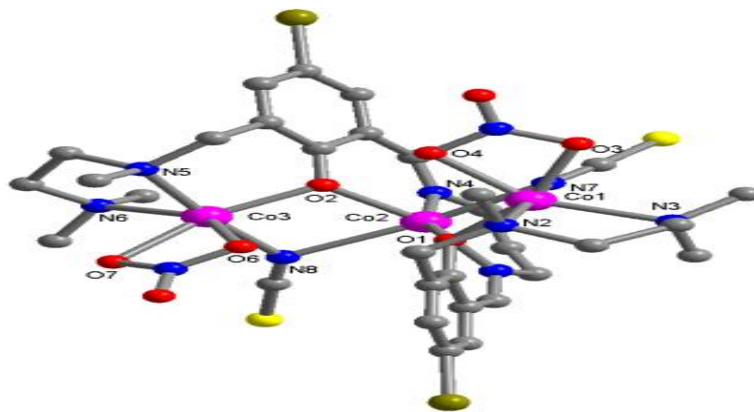


Figure 4: the tri-nuclear complex 5 [34]

2.4. [1,3-diphenyl-3-(phenylamino)propan-1-one (1)]

Hf (OTf)₄ was discovered to be a very effective catalyst for the Mannich process. Under solvent-free conditions, as little as (0.1-0.5) mol percent Hf (OTf)₄ may catalyze the high-yielding synthesis of a variety of aryl and alkyl ketone-based Mannich bases. The inclusion of Hf (OTf)₄ resulted in good region- and diastereo-selectivity in the synthesis of alkyl ketone-based Mannich bases. The ¹H-NMR tracing of the H/D exchange reactions of cyclopentanone and acetophenone in MeOH-d₄ revealed that the coordination of Hf (OTf)₄ with ketone might permit its quick keto-enol tautomerization, hence contributing to the overall promotion of the Mannich reaction. [35], see Figure 5.

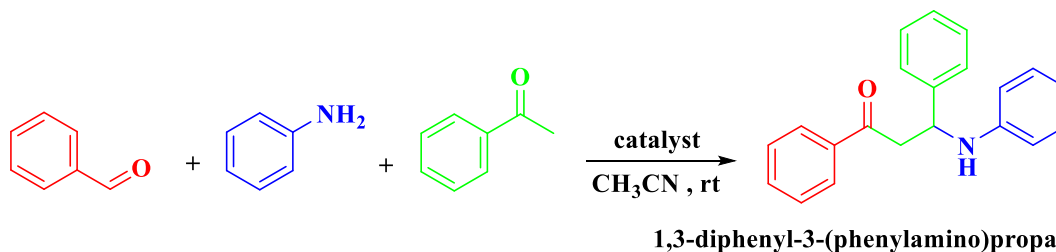


Figure 5: the catalytic effect on Mannich reaction [35].

2.5. Mannich Bases and Bis-(Mannich bases) of Isoindolin-1,3-dione (1)

Synthesized novel Mannich bases compounds through reaction of formalin with the appropriate amine or di-amine and iso indolin-1,3- di one (1). By using dihydrazide or hydrazide as amino compounds in the reaction of Mannich base with 1, both (bis-Mannich and Mannich) bases were synthesized. The preparation of Mannich bases involved the incorporation of 1, showcasing the compositional potential of sec-Mannich bases is described as a precursor. The ketonic Mannich bases with N-alkylation of 1 was examined [36], see in Figure 6.

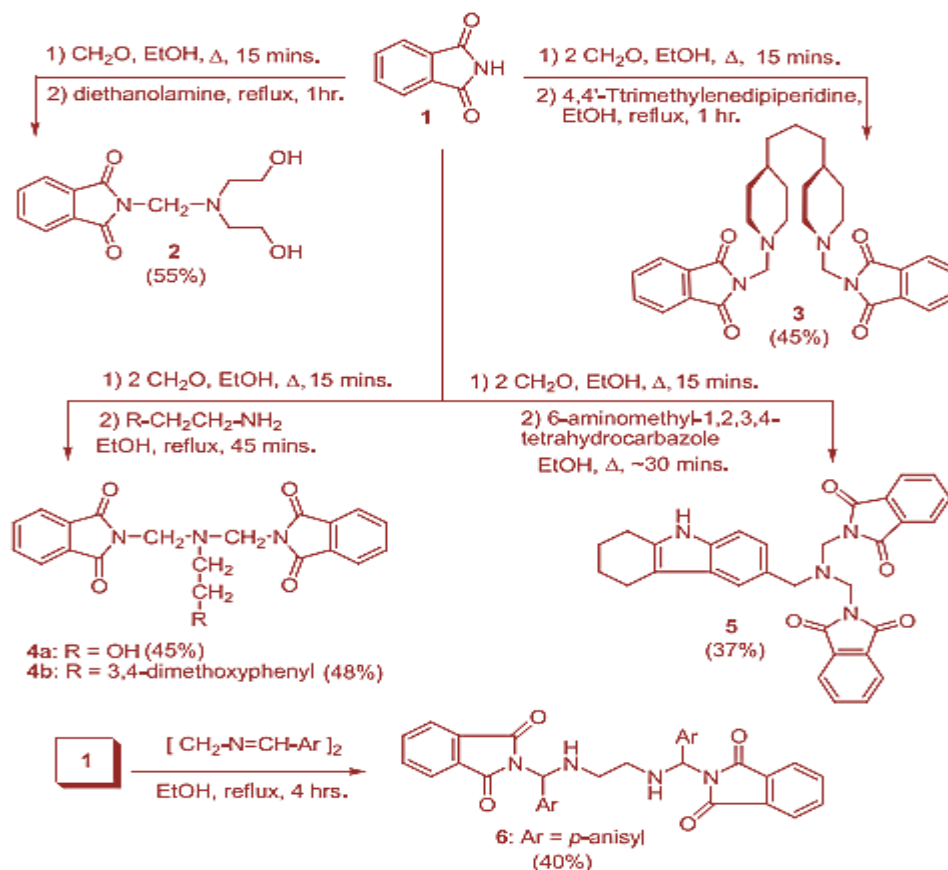


Figure 6: Mannich bases and compound (1) [36].

3. Versatile applications of Mannich bases

3.1. Catalytic properties

3.1.1. The catalytic properties of complexes for the ring-opening polymerization of ϵ -caprolactone

Synthesized a chain of Mannich base ligands HA = [2, 4 - *t* Bu₂ -6-CH₂ N Me₂ -PhOH, HB= 2, 4-*t* Bu₂ -6-CH₂NEt₂-PhOH (HB), HC= 2, 4-*t*Bu₂-6-CH₂ Py- Ph OH, HD= 2-*t*Bu-4-Me-6-CH₂ Py- Ph OH, HE= 4-*t*Bu- 2, 6-(CH₂Py)₂-Ph OH] through (Mannich reaction) by used form aldehyde and phenol reaction with amine (secondary). A chain of The ligands' amine elimination interactions with Ln[N(SiMe₃)₂] guided in the formation of homo-leptic complexes [LaA₃ (1), LaB₃ (3), GdA₃ (2), GdB₃ (4), GdC₃ (6), LaC₃ (5), LaD₃ (7), GdE₃ (10), Gd D₃ (8) and LaE₃ (9)]. NMR spectra were used to describe Complexes 1, 3, 5, 7, and 9, and single-crystal X-ray diffraction was included to identify the structures of Complexes 3 and 5. Complexes 3 and 5 has same structure, with distorted octahedral geometry in the lanthanum core. Under this geometry, O₍₁₎, O₍₂₎, and O₍₃₎ atoms occupied two platforms, although N₍₁₎, N₍₂₎, and The NMR spectra of a -caprolactone oligomer indicated that the ROP of -caprolactone occurred via a coordination-insertion process. These Mannich base monophenolate lanthanide complexes' catalytic characteristics for ROP of -caprolactone have been investigated. For -caprolactone polymerization, all complexes show strong catalytic activity. The NMR spectra of a -caprolactone oligomer indicated that the ROP of -caprolactone occurred via a coordination-insertion process. N₍₃₎ atoms occupied the remaining three spaces. All complexes were analyzed by integratungelemental analyses and infrared spectroscopy. The catalytic properties of these Mannich base monophenolate lanthanide complexes for ROP of ϵ -caprolactone have been studied. All complexes showed high catalytic activity for ϵ -caprolactone polymerization. The NMR spectra of a -caprolactone oligomer indicated that the ROP of -caprolactone occurred via a coordination-insertion process. These Mannich base monophenolate lanthanide complexes' catalytic characteristics for ROP of -caprolactone have been investigated. For -caprolactone polymerization, all complexes show strong catalytic activity. The NMR spectra of a -caprolactone oligomer indicated that the ROP of -caprolactone occurred via a coordination-insertion process. [37], see Figure 7.

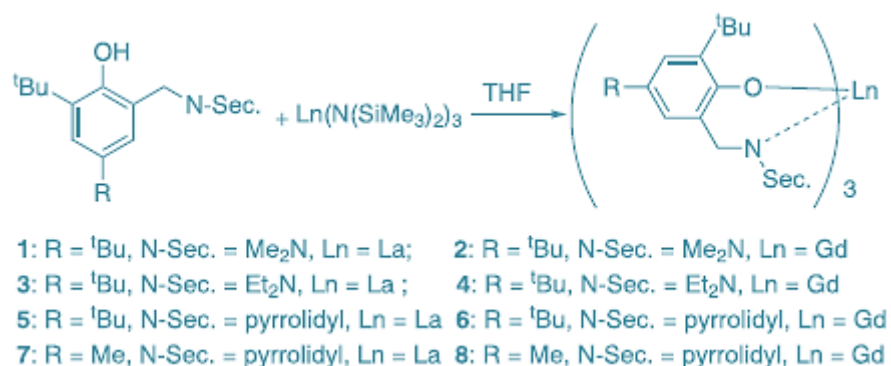


Figure 7: synthesis of complexes 1–8 [37]

3.1.2. The Trinuclear Copper Complexes as Catalysts in Oxidation Reactions.

In 2018, Avijit et al., [38] successfully synthesized three ligands, namely The ligands used in this study, Namely H₂L₁, H₂L₂ and H₂L₃ where N, N – bis (3, 5-dimethyl-2-hydroxy benzyl)- N', N' -dimethyl-1,3-diamino propane, H₂L₁; N, N -bis(3,5-dimethyl-2-hydroxy benzyl)-N', N'-dimethyl-1,2-diaminoethane], H₂L₂ and N,N-bis(3,5-dimethyl-2-hydroxy benzyl)-N', N'-diethyl-1,2-diaminoethane, H₂L₃ were reacted with [Cu.(CH₃COO)₂.H₂O] to form dinuclear compounds [Cu₂L₂] (1-3). Subsequent reactions of these dimeric compounds with [Cu (ClO₄)₂.6H₂O], and [NaN₃] resulted in the formation of trinuclear compounds, [(CuL₁)₂(μ_{1,1}-N₃)₂.Cu(H₂O)].CH₃OH (1), [(CuL₂)₂(μ_{1,1}-N₃)₂.Cu(H₂O)].2CH₃OH (2), and [(CuL₃)₂(μ_{1,1}-N₃)₂.Cu(H₂O)].CH₃OH (3), in a 1 to 1 manner. The structures of these complexes were characterized using X-ray diffraction and CHNS analysis. All complexes exhibited catecholase-like and phenoxazine synthase-like activities in the presence of 3, 5 – di – tert-butyl catechol and aminophenol, respectively. The catalytic oxidation reactions resulted in the production of H₂O and H₂O₂ during the catecholase - like and phenoxazine synthase activities respectively. These findings demonstrate the potential of these trinuclear copper complexes as catalysts in oxidation reactions and highlight their promising applications in various catalytic processes, see Figure 8.

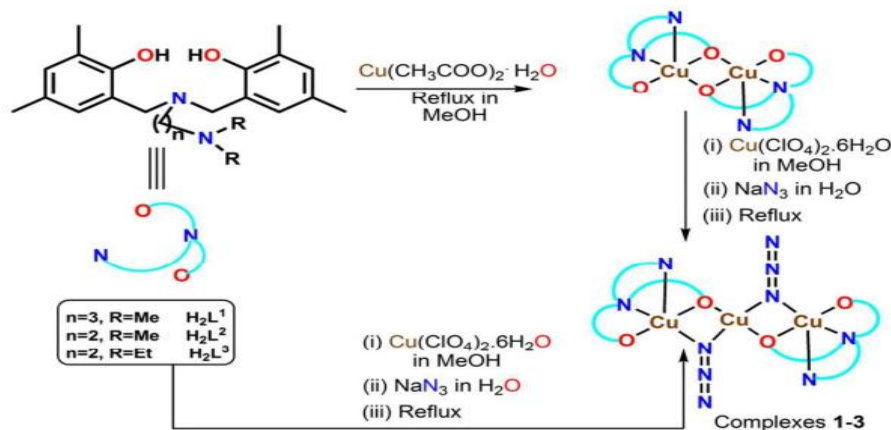


Figure 8: syntheses of complexes 1-3

3.2. Biological Activity

3.2.1. The anti-microbial Activity of Prepared Compounds Towards Selective Bacteria and Fungi

In 2016 Rehab et al., [39] synthesized novel compound [N-(4- morpholino methyl)-1,8–naphthal imide] =Mannich base ligand, reacted with metallic ion Pd²⁺, Pt⁴⁺, Ru³⁺, Rh³⁺, and were diagnosis through FTIR, UV-Vis and ¹HNMR spectroscopy, A.A., C.H.N.S, as well as conductance and magnetic measurements. The results display for all complexes were create in [Metal: Ligand] [1: 2] mole ratio but for Palladium²⁺ complex which was create in [Metal: Ligand] [1:1] mole ratio. The anti-microbial action of prepared compounds has been tested towards some selective bacteria, for example P. aeruginosa), B. Subtilis and some selective fungi C. albicans), A. flavus through diffusion method [39], see Figure 9.

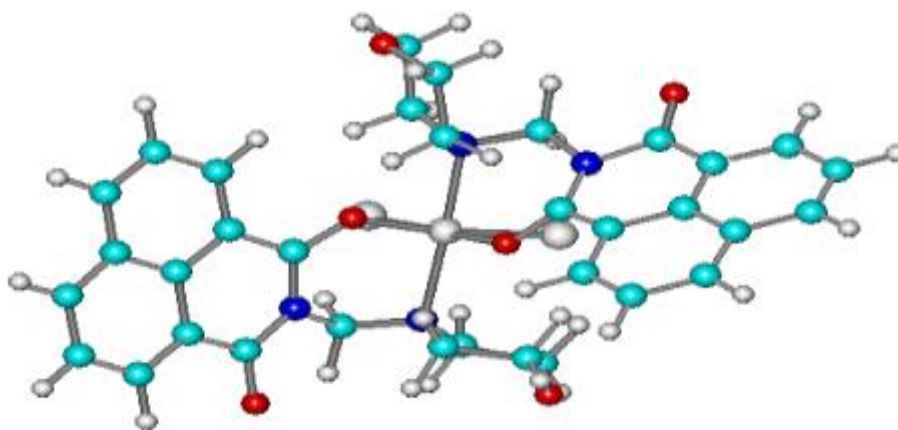


Figure 9: synthesis of [RuL₂] complex [39].

3.2.2. The Anti-microbial Activity of Prepared Compounds Towards Selective Bacteria and Fungi

Mannich bases been synthesized from (TBB)= N-(phenyl(thiomor pholino) methyl) benzamide, (TBA)= N-(phenyl(thiomorpholino)methyl) acetamide, (TBC)= 1-N-(phenyl (thiomorpholino)methyl) carb amide and (TBN)= N-(phenyl (thiomor pholino) methyl) nicotin amide. TLC, FTIR, MS Spectra, ¹H, ¹³C-NMR and C.H.N.S. The antibacterial and antifungal activity of the prepared compounds was tested against specific types of bacteria and fungi. (TBC) show very good efficiency towards *S. aureus*, *P. aeruginosa* and *E. coli*; (TBA) show very good efficiency towards *P. species*, *A. niger* and *C. albicans*, compared to other Mannich bases [40], see Figure 10.

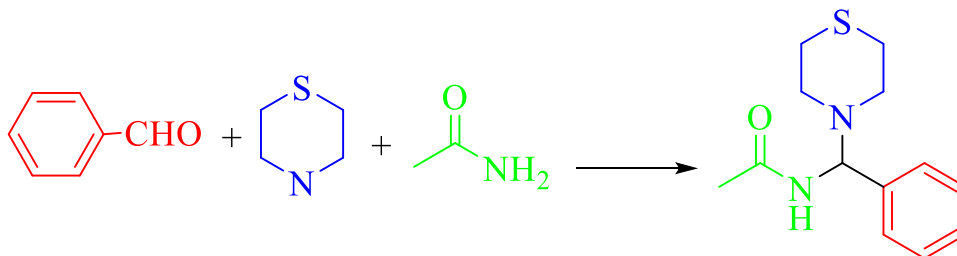


Figure 10: synthesis of (TBA) [40].

3.2.3. The Anti-bacterial Activity of Prepared Compounds

Prepared a novel ligand BTBPBO = [3-4-[4-(Benzo thiazol-2-yl-amino)-benzyl]-phenyl iminobutan-2-one oxime] from the reaction of 4,4-dia aniline methane with 2-mercapto benzo thiazole, then the produce of this period reacted with diacetylmonoxime. The compounds were organized from reaction of (BTBPBO) with metal ion (ii) [Co, Cu, Cd and Hg] and diagnosis through the spectral and physical ways to govern the stereo-geometries of prepared compounds, which showed that it has a shape tetra hedral but Co(ii), Cu(ii) compounds were octahedral. The prepared compounds display good anti-bacterial efficiency [41], see Figure 11.

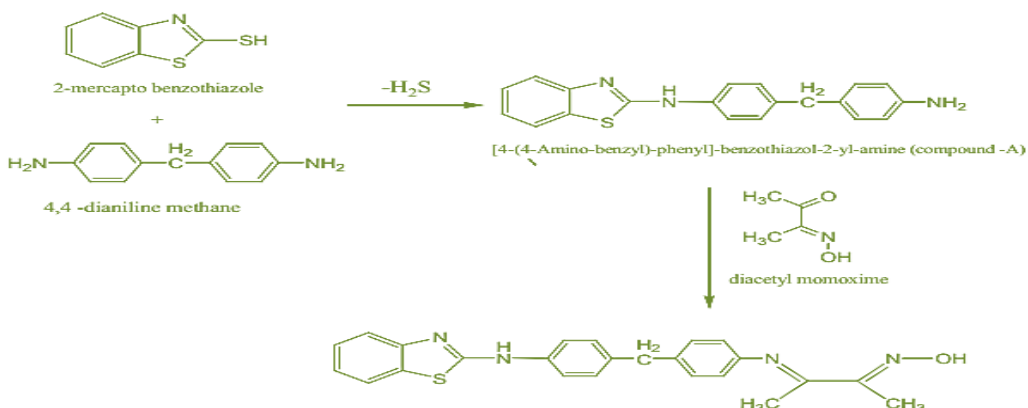


Figure 11: synthesis of Mannich bases [41].

3.2.4. Anti-bacterial Activity of Compounds Towards Several Choices Bacterial

Novel chain compounds of mixed ligands contained (La)= 1,10-phenanthroline and Mannich ligand (Lb), where Lb= The compounds Lb₁, Lb₂, and Lb₃ were synthesized as N-[1-Morpholino-(3-methoxy-4-hydroxy)benzyl] acetamide (M.B.A); N- [1-Morpholino-(3-methoxy-4-hydroxy) benzyl] urea (M. B.U) and N-[1-Morpholino-(3-methoxy-4-hydroxy) benzyl] benzamide (M.B.B) respectively with Fe(II), Ni(II), Co(II), Zn(II) and Cu(II) ions. General formula with preparing complexes: [M (Lb)₂ (La)] n H₂O and octahedral structure. The compounds were diagnosis by UV/Vis, conductance, FT/IR, and CHN as well as magnetic data. Through the results of the measurements, the ligands showed a bidentate coordination with the metal ions by O /atom for the CO-NH₂ group, and N/atom for the C₄H₉NO group, while the ligand La is coordinated through the two nitrogen atoms. Anti-bacterial actions of compounds were studies towards several choice bacterial [42], see Figure 12.

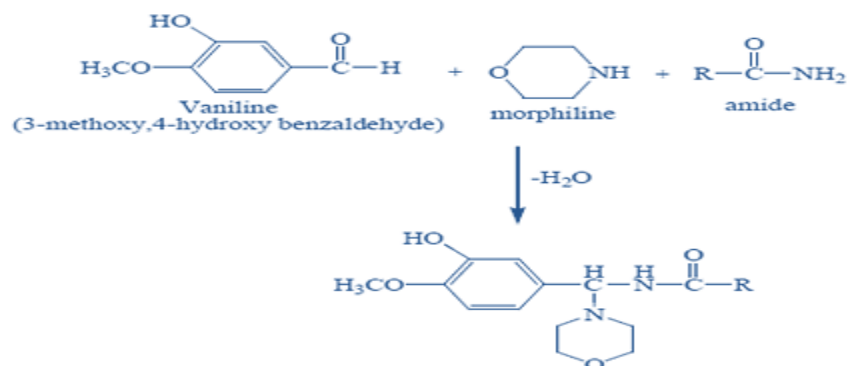


Figure 12: preparation of ligand [42].

3.2.4. The Anti-microbial Activity of Mannich Bases

Reacted aldehyde and produce primary aliphatic amines with (5-methyl -1Hs -triazole-3- thiol)(1) to generate the Schiff-Mannich base [2-methyl-6-substituted-6,7-dihydro-5H-s-triazolo[5,1-b]-1,3,5-thiadiazines] (cyclic), When primary aromatic amines were coupled with 3-methyl-1-((substituted-amino)methyl)-1H-s-triazole-5-thio1 (8) in the laboratory or boiling ethano1. Without integrating the secondary amine, this reaction produced 3-methyl-5,6-dihydro-s-triazolo[3,4-b]-1,3,4-thiadiazol. These reactions followed the Mannich reaction criteria. Furthermore, Schiff's bases (unicycles) were synthesized by reacting aromatic aldehydes with compound (8) under varied circumstances. All unique compounds' structures were proven by C.H.N.S. In addition to the anti-microbial activity, some prepared derivatives showed a high ability to remove certain ions (Pb(ii), Mg(ii), Ca(ii) and Cd(ii)) from aqueous solutions [43], see Figure 13.



Figure 13: the potential mechanism for the removal of metal and heavy metal ions [43].

3.2.4. The Anti-bacterial Activity Testing Against Specific Types of Bacteria

Reacted salicylaldehyde and 1,3-cyclohexanedione with urea to prepare Mannich base, 1-((2,6- dioxo cyclohexyl) (2-hydroxy1phenyl) methyl) urea and through its reaction with metal ions a group of complexes (1a-1e) were prepared, where the prepared compounds were identified by different spectroscopic methods EPR, UV-visible,

FT-IR, ^1H - ^{13}C NMR and from the results of the measurements the octahedral geometry of complexes was determined. Using the diffusion method. The prepared compounds were subjected to biological activity testing against specific types of bacteria *S. aureus*, *P. aeruginosa*, *K. pneumoniae* and *E. coli* as the results showed that the effectiveness of the complexes was higher than the reference (Ciprofloxacin) used [44], see Figure 14.

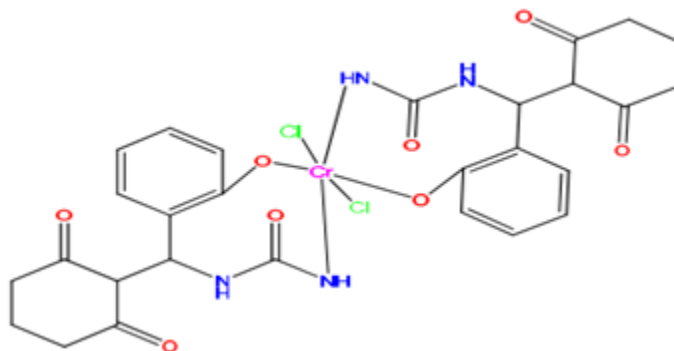


Figure 14: structure of Cr complex [44].

3.2.5. The anti-bacterial Activity of the Prepared Compounds

Prepared Ni(II), Cd(II), Co(II), Cu(II) and Fe(III) complexes derived from tri-azole new tri-dentate ligands and diagnosis using various spectroscopic ways. As proposed by the conductivity data, all the compounds were of an electrolytic nature. Ligand trizo1 is coordinated with the metal ions in the complexes through an atom (N) in a group (amino) and an atom (S) in a group (thiophenolic) and this is confirmed by the measurements of (FTIR). The thermal stability of the prepared compounds was studied, which was found to have an octahedral geometry. When testing the bacterial activity of the prepared compounds, it was found that the compounds containing ions (Nickel(ii), Cobalt(ii), and Copper(ii)) as a metal centre showed good results against the selected bacterium types (*B. subtilis*, *S. pyogenes* and *P. vulgaris*) compared to the standard [45], see Figure 15.

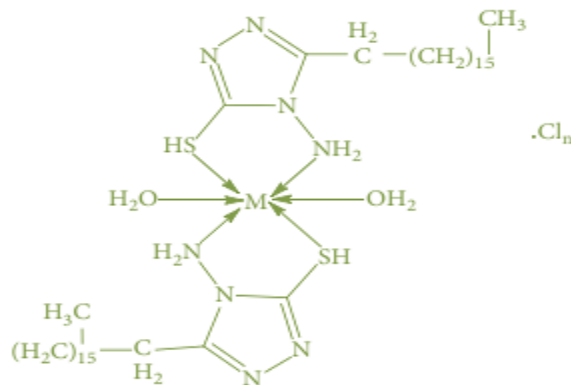


Figure 15: proposed structure of complexes [45].

3.2.6. The anti-bacterial Activity of Ligand and Complexes

Reacted 4-N, N-dimethyl benzaldehyde A Schiff-Mannich base, 5-((1H-indol-1-yl)methyl thio)-N-(4-(dimethyl amino)benzylidene)-1,3,4-thiadiazol-2-amine (L), was synthesized by combining 2-amino-5-mercapto-1,3,4-thiadiazole with indole. This ligand, L, was then reacted with various metal ions including Co(II), Pd(II), Ni(II), Cu(II), Au(III), and Pt(IV) to form complexes. Using different spectroscopic techniques such as ^1H - ^{13}C /NMR, UV/Vis, CHNS, and FT-IR, complexes have been identified. Based on the measurement results, it was determined that the ligand L exhibited a tridentate behavior when interacting with the metal ions. The antibacterial activity of the prepared compounds was evaluated against specific strains including *Klebsiella pneumoniae* (G-), *Bacillus subtilis*, and *Staphylococcus aureus* (G+). The results indicated that the complexes demonstrated higher antibacterial activity compared to the ligand alone. Furthermore, the HyperChem 8.0.7 program was utilized to study the compounds in the gas phase and calculate the binding energy, HOMO, LUMO, heat of formation, infrared vibration of the bond, and the electrostatic potential [46], see Figure 16.

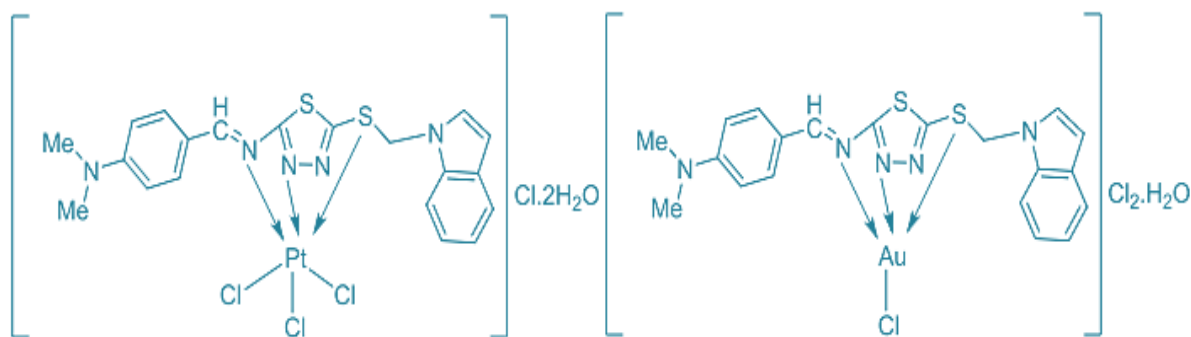


Figure 16: structure of complexes (Pt and Au) [46].

3.3. Corrosion inhibition

3.3.1. The effect Mannich Base as Corrosion Inhibition on Copper in 1 M HCl Medium

Synthesized a Mannich base= MFC= N-(4-(Morpholino methyl Carbamoyl Phenyl) furan-2-carbox amide and diagnosis through ^1H , ^{13}C -NMR and FT-IR. By measuring electrochemical impedance spectroscopy (EIS), Periodic voltage measurement, weight loss wt., and active dynamic polarization (CV), the effect of MFC inhibition on copper in 1 M HCl medium was studied. To describe the corrosion inhibitor mechanism, thermo dynamic parameters, for example, enthalpy, free energy and entropy were personalized. With increasing Adding an inhibitor to the copper metal controls this copper oxidation, as revealed in the CV. Significant morphological improvement over the copper surface with the addition of the inhibitor. This was demonstrated by surface analysis using SEM [47], see Figure 17.

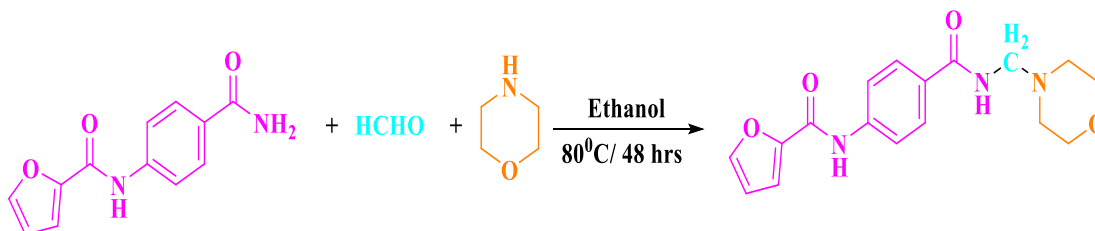


Figure 17: Synthesis of MFC [47].

3.3.2. Mannich-base H.M.P as a corrosion inhibitor for mild steel in a 0.2 M

A new Mannich base called 2-((4-(2-hydroxy-4-methylbenzyl)-piperazin-1-yl)methyl)-5-methylphenol (H.M.P) was successfully synthesized. The compound's characterization was carried out using various spectroscopic techniques such as C.H.N.S, UV-Vis, ^1H , and ^{13}C -NMR. To assess its effectiveness as a corrosion inhibitor for mild steel in a 0.2 M A substance has been evaluated for weight loss and potential dynamic polarization methods in an H_2SO_4 solution. The results revealed that the compound's inhibitory activity increased with increasing inhibitor concentrations. However, the inhibitory impact faded within 96 hours. Analysis of potential dynamic polarization data revealed that the inhibitor acted as a mixed-type corrosion inhibitor; density functional theory (D.F.T) calculations were performed to gain a more accurate recognizing of the relationship between the molecular structure of H.M.P and its inhibitory properties. This research sheds light on the synthesis, characterization, and prospective use of the new Mannich-base H.M.P as a mild steel corrosion inhibitor in acidic situations. The findings on its inhibitory properties and the relationship with its molecular structure offer important insights for future research and the development of effective corrosion inhibitors for various industrial applications [48]; see Figure 18.

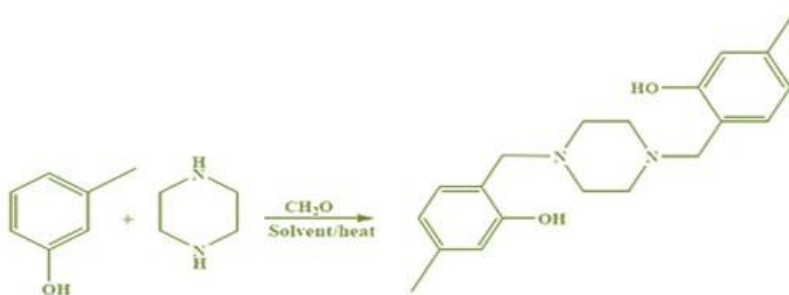


Figure 18: Synthesis of (HMP) [48].

3.4. Antioxidant activity

3.4.1. Antioxidant activity of Mannich-bases

Mannich-bases for cycloalane derivatives were synthesized, and their antioxidant activity was determined using the D.P.P.H free radical-scavenging assay. Via $^1\text{H}/\text{NMR}$, FT/IR, $^{13}\text{C}/\text{NMR}$ and MS spectroscopic result, the composition of the compounds was confirmed. The highest antioxidant activity was shown by the complex de Maniche derivative of cyclofalone with diethylamine (2a), with an IC_{50} value of 39.0 M. According to structure-activity connection research, compounds suffering a higher pKa value for the Mannich base have more activity, as proven by lower IC_{50} values [49], see Figure 19.

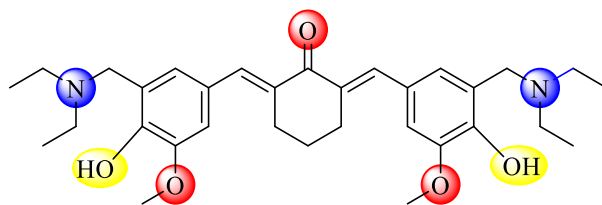


Figure 19: ^1H -NMR spectrum of compound 2a [49].

3.4.2. Antioxidant properties of Mannich bases

Three-part Mannich reaction involving benzaldehyde, either acetophenone or 4-iodoacetophenone, and various substituted anilines, catalysed by di-ethanol ammonium chloroacetate. The Mannich bases (MBs) obtained from this reaction showed excellent yields. All compounds were comprehensively characterized using C.H.N.S, ^1H , ^{13}C -NMR, UV-Vis, and FT-IR spectral analyses. The *in vitro* antioxidative potential of the synthesized MBs was evaluated utilized in this study, the antioxidative activity of compounds containing anisidine moiety was investigated using the Thermodynamic investigation of 2,2-diphenyl-1-picrylhydrazyl radical and density functional theory (D.F.T). The results illustrated that these compounds have slight antioxidative implementation. Furthermore, getting rid of iminium ions is required to eliminate water, which was suggested as a key rate-determining step for the overall reaction process. This study provides valuable insights into the synthetic efficiency and antioxidative properties of the produced MBs, offering potential applications in medicinal chemistry and other fields requiring antioxidant agents. The use of diethanol ammonium chloroacetate as a catalyst represents a significant advancement in the Mannich reaction, facilitating the synthesis of diverse MBs with promising biological activities [50]; see Figure 20.

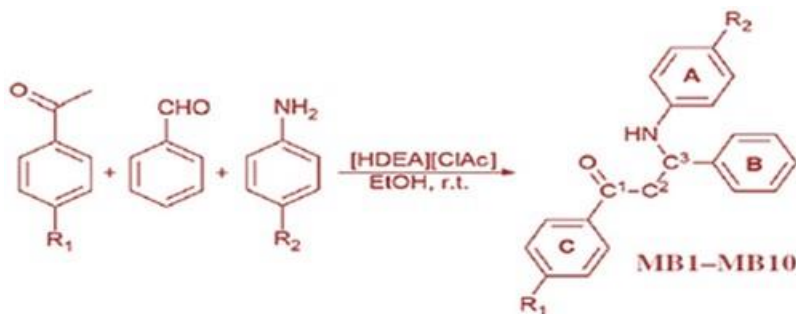


Figure 20: Structure of Mannich bases [50].

3.4.3. Antioxidative Activity of Mixed Ligand Complexes

Synthesized a novel series of mixed ligand complexes containing Coordinated with Co(II), Ni(II), Co(II), and Zn(II) ions, 8-hydroxyquinoline (Q) and a bis-phenolic Mannich base, 2,2'-((2-(dimethyl-amino)ethyl)-azanediyl)-bis(methylene)-bis(4-(tert-butyl))bis(4-(tert-butyl))bis(methylene)-bis(4-(tert-butyl))bis(phenol), for med complexes. The compounds were characterized using magnetic measurements, spectral studies, conductance measurements, and (CV) cyclic voltammetry. Magnetic and spectral data revealed an octa hederal structure for all the complexes. The redox behaviour of metal complexes was investigated through cyclic voltammetry. Using the diffusion method, the synthesized compounds were tested for their anti-bacterial activity against *S. aureus* and *E. coli* their anti-fungal activity against *A Niger* and *Penicilium* species. Remarkably, all the compounds exhibited potent antimicrobial properties against the tested fungi and bacteria. Moreover, the compounds demonstrated significant antioxidant capacity based on antioxidant assays. This indicates their potential as effective antioxidants, which is of great interest for various applications in pharmaceuticals, food preservation, and other fields that require antioxidative properties. The introduction of a bisphenolic Mannich base in combination with 8-hydroxyquinoline resulted in the formation of promising mixed ligand complexes, showing both strong anti-microbial activity and antioxidative potential. This study contributes valuable information for developing novel complexes with potential applications in medicine and other fields [51], seen Figure 21.

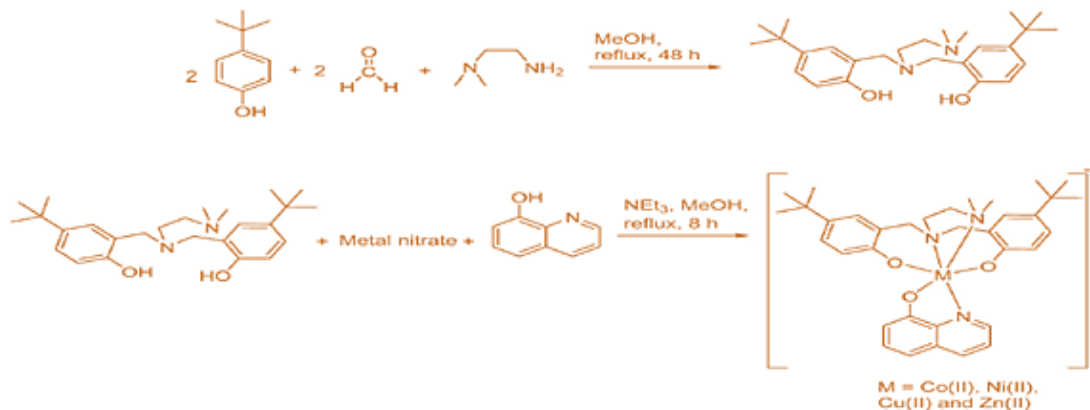


Figure 21: synthesis of new compounds [51].

3.5. Anti-urease Activity of Mannich Bases

Prepared Mannich base {MMP= 1-[(4-methoxy phenyl)(2-methylidene cyclohexyl) methyl]pyrrolidine through reaction (4-methoxy benzaldehyde, cyclohexanone and pyrrolidine). The MMP reaction with chloride salt (ii) (Cu, Ni, Co, and Fe) gives their complexes. The prepared chemicals were ascertained utilizing FT/IR, TGA, ¹H/NMR and ¹³C/NMR measurement, and their potential as anti-urease agents was evaluated. The Mannich base exhibited with an IC-50 value of 224.83, 0.84 M, it has modest anti-urease action. Current research can give an intuitive approach for synthesizing a Mannich base with efficient anti-urease scaffolds [52], see Figure 22.

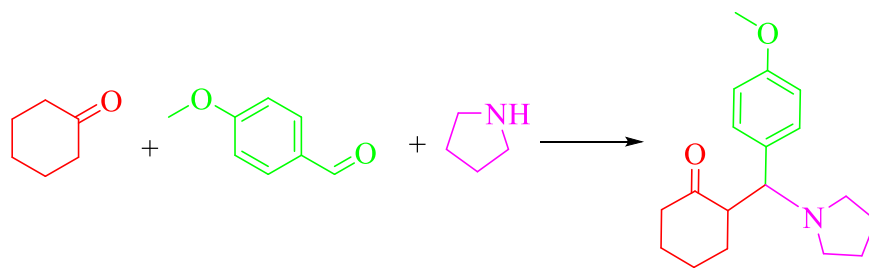


Figure 22: Synthesis of Mannich base (MMP) [52].

3.6. Anticancer activity

3.6.1. Anticancer activity of Mannich bases

Four times higher activity than standard tamoxifen, Mannich's-base compound N-[(diphenyl amino) methyl] acetamide has been shown to be used to overcome drug resistance to estrogen receptor protein. Using Accelrys Discovery Studio version 2.5 client software, the X-ray crystal structure of the estrogen receptor protein (PDB ID 2YAT) was downloaded from the Protein Data Bank (PDB) and docked with the target Mannich bases. The best compound was also identified, and its anticancer activity was confirmed using in vitro MTS analysis using Raju and Jurkat cell lines based on the results of the in silico analysis of the target compounds with standard tamoxifen [53], see Figure 23.

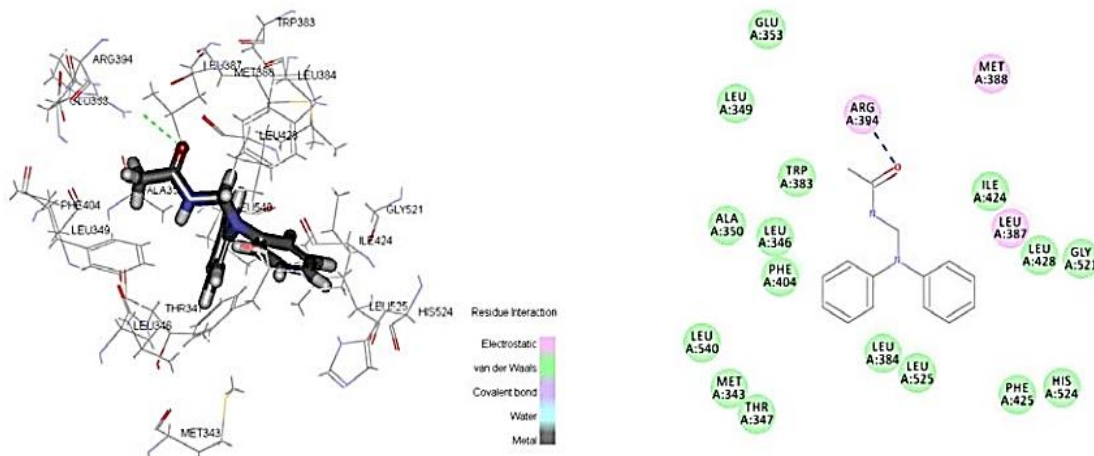


Figure 23: Association of N-[(diphenylamino)methyl] acetamide with amino acids [53].

3.6.2. The anti-cancer Activity of Compounds

Synthesized novel aminonaphthaquinones derived from law sone using Mannich reaction, at lab temperature for anti-cancer use. Munch bases compounds have been prepared, including various amines and aromatic aldehydes with a product yield ranging from medium to high. The anti-cancer efficiency (MTT) with nuclear morphology assessment (DAPI staining or 40, 6-diamidino-2-phenyl indole), apoptosis assessment Tests such as ethidium bromide staining or acridine orange, hemo lysis, and DNA ladder assay can be conducted on the human liver carcinoma cell line Hep-G2 to obtain results. Additionally, one of the basic compounds derived from dried henna (*Lawsonia inermis*) is available leaves is Lawsone (2-hydroxy1-1,4-naphtha quinone) [54], see Figure 24.

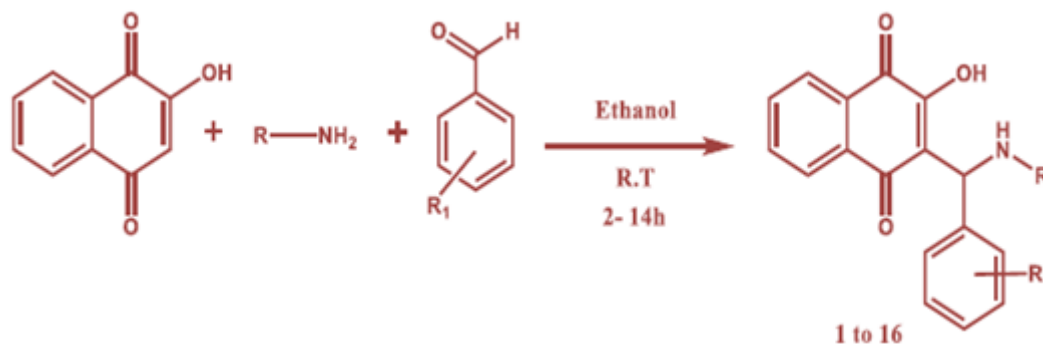


Figure 24: Production of Mannich base compounds [54].

3.6.3. The anti-cancer Activity of Ligand and both Complexes

Synthesized new ligand Mannich base containing the 1,3,4-oxadiazole ring was synthesized and characterized using FT/IR, UV, elemental analysis, $^1\text{H}/\text{NMR}$, and $^{13}\text{C}/\text{NMR}$ techniques. Subsequently, a complex of Pt+4 and Cu+2 was prepared by complexation with the synthesized ligand and characterized using UV, FT-IR, flame atomic absorption, magnetic susceptibility, and molar conductivity. The results confirmed the successful synthesis procedure. The antibacterial activity and cytotoxicity. The attributes of the ligand and its metal complexes were investigated. Antibacterial and antifungal activity has been illustrated, and identified against *E. coli* (gram-negative), *S. aureus* (gram-positive), *C. albicans*, and *A. flavus*. Cytotoxicity was examined on a normal cell line (MDCK), and the findings revealed that the LD_{50} of L reached 936 g/ml, the toxicity of Pt reached 875 g/ml, and the toxicity of Cu was 1356 g/ml. The ligand and both complexes had been further evaluated for toxicity against a human lung cancer cell line, and the findings disclosed that the LD_{50} of the Pt complex at 615 g/ml. toxicity of Cu complex= 653 $\mu\text{g}/\text{ml}$ and toxicity of free ligand= 620 $\mu\text{g}/\text{ml}$. Low MIC was resulted for the ligand and both complexes compared with Amoxicillin and Ampicillin as antibiotic [55], as seen Figure 25.

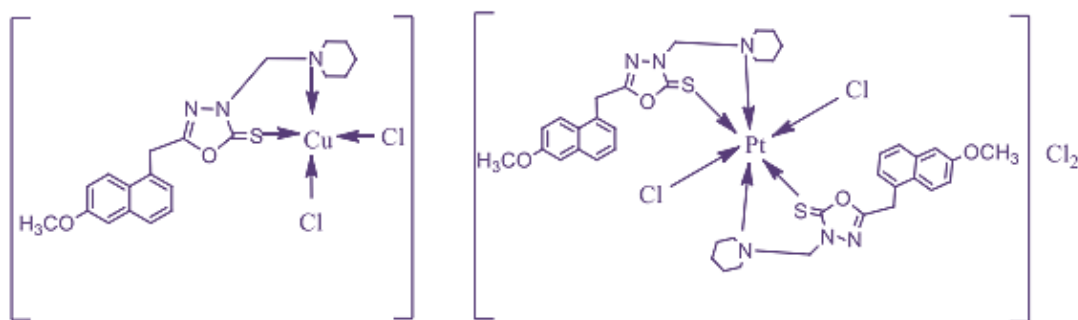


Figure 25: flowchart for the synthesis of ligands and complexes [55].

3.6.4. The anti-cancer Activity of the Prepared Compounds

Prepared ligand Mannich base [3-((3-(4-hydroxyl phenyl)-5-mercapto-4H-1, 2, 4-triazol-4-yl)amino) benzo furan-2(3H)-one] (L), which is derived from acetone and formaldehyde with secondary amine (heterocyclic compound possesses two rings), which in turn interacted with the metal ions Ni(II), Cu(II), Au(III) and Pd(II) to prepare the complexes, where the prepared compounds were identified by different spectroscopic methods C.H.N.S, UV-Vis, TGA and molar conductivity, flame atomic absorption and magnetic susceptibility from the results of the measurements, the square-planar geometric structure was suggested for the metal complexes with chemical formulas $[\text{PdLCl}]\text{Cl}_2 \cdot 2\text{H}_2\text{O}$, $[\text{NiLCl}]\text{Cl}_2 \cdot 6\text{H}_2\text{O}$, $[\text{AuLCl}]\text{Cl}_2 \cdot 3\text{H}_2\text{O}$ and $[\text{CuLCl}]\text{Cl}_2 \cdot \text{H}_2\text{O}$. Using the MMT method, a toxicity examination of the prepared compounds was carried out on the brain cancer cell line (AMJM), where it was found that the gold complex is the highest toxicity among the tested compounds [56], see Figure 26.

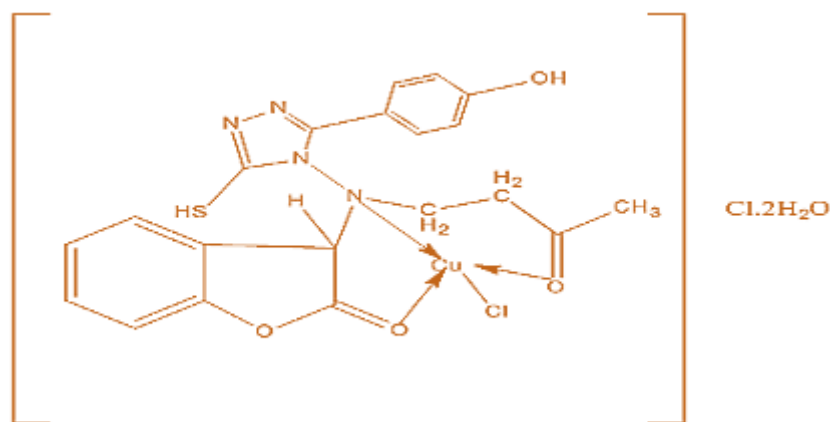


Figure 26: Proposed structure of complexes [56].

3.7. Mannich base Ligand as a Fluorescence Sensor

The complexes were produced by the ligand H₂L (N, N'-di methyl -N, N'-bis (2- hydroxyl-3-methoxy-5-methyl benzyl) ethylene diamine) responded with Zn (ii) ion in the presence of (bridging co - ligands), azide/acetate or thio cyanate/acetate, chloride. (di-nuclear), [Zn₂L(SCN)₂(H₂O)] · H₂O (1), [Zn₂L(N₃)(CH₃CO₂)₂] (2) and [Zn₂LCl₂(H₂O)] (3), while Cd(ii) ion produced complexes (tetra-nuclear), [Cd₄L₂(N₃)₂(CH₃CO₂)₂] · 3H₂O (4), [Cd₄L₂(SCN)₂(CH₃CO₂)₂] · 2H₂O (5), [Cd₄L₂Cl₄] · H₂O (6) with (H₂L) and co-ligands. The possibility of using the ligand (H₂L) as a fluorescence sensor that automatically detects the zinc ion ⁽⁺²⁾ is very effective in the work, as (7.69) nM represents the detection limit for the said ion. it is greater than most of the fluorescence sensors reported for this ion with the correlation constant K = 1.508 × 10¹⁰ M⁻² zinc complex ⁽⁺²⁾ was also used as a sensor. Recently, zinc complexes may be good as a picric acid sensor because of their high sensitivity, structural simplicity, and cost calculation [57], Figure 27.

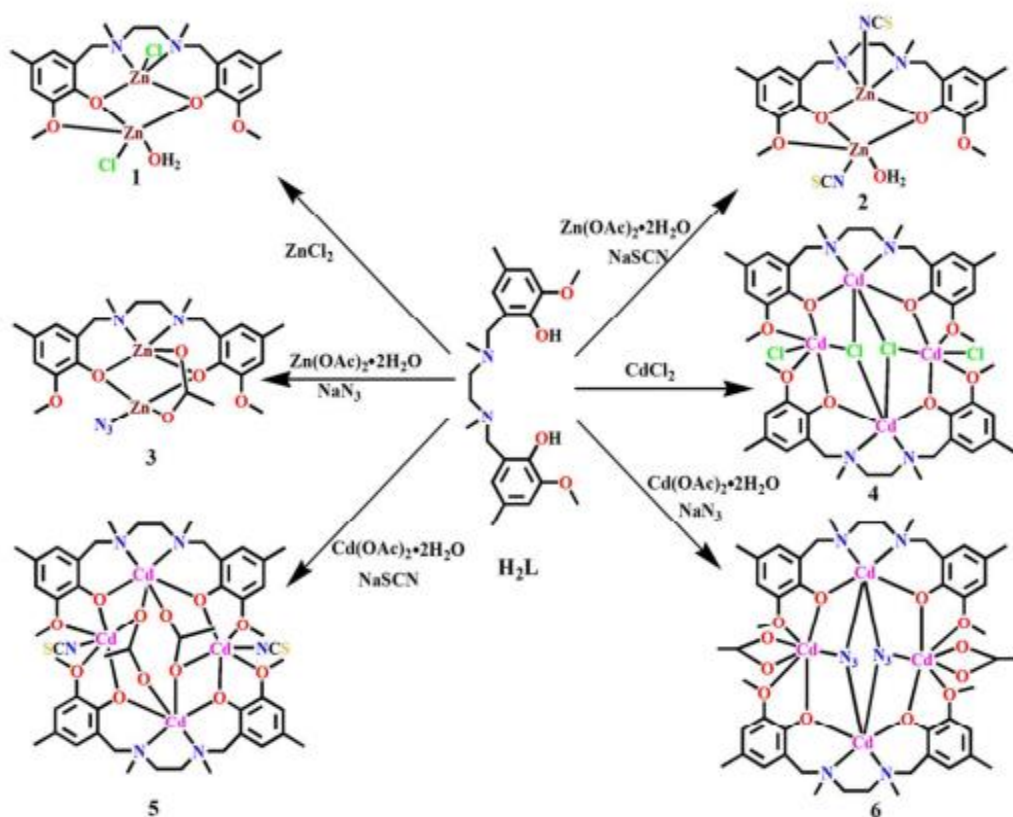


Figure 27: Synthesis of complexes 1-6 [57].

3.8. The potential Applications of Mannich-base in various fields, including Magnetism and Materials Science

Synthesized metal complexes by reacting the Mannich-base ligand H₂L where N, N'-di methyl -N, N' -bis (2 - hydroxy -3- methoxy- 5-methyl benzyl) - ethylene - di amine with the metal ions Co and Ln, where Ln= Dy, Tb, and Gd, resulting in three complexes: Co₂Gd₂ L₂(μ₄-CO₃)₂(NO₃)₂] (1), [Co₂Dy₂L₂ (μ₄-CO₃)₂ (NO₃)₂] (3) and [Co₂Tb₂L₂(μ₄-CO₃)₂(NO₃)₂] (2) and Measurement results revealed that the geometric structure around the cobalt ions in all complexes is a distorted square pyramid. In contrast, the lanthanum ions exhibit; whole the complexes exhibited a spherical capped square antiprism and spherical tricapped trigonal prism geometry. The magnetization measurements showed a pronounced magnetization relaxation in complexes 2 and 3, whereas the effect was almost negligible in complex 1. This study highlights the distinct geometric arrangements and magnetization behavior of the synthesized metal complexes, providing valuable insights into their structural and magnetic properties. The findings our team in comprehending the coordination chemistry of various metal-ligand interactions, which may have implications for their potential applications in various fields, including magnetism and materials science [58]; see Figure 28.

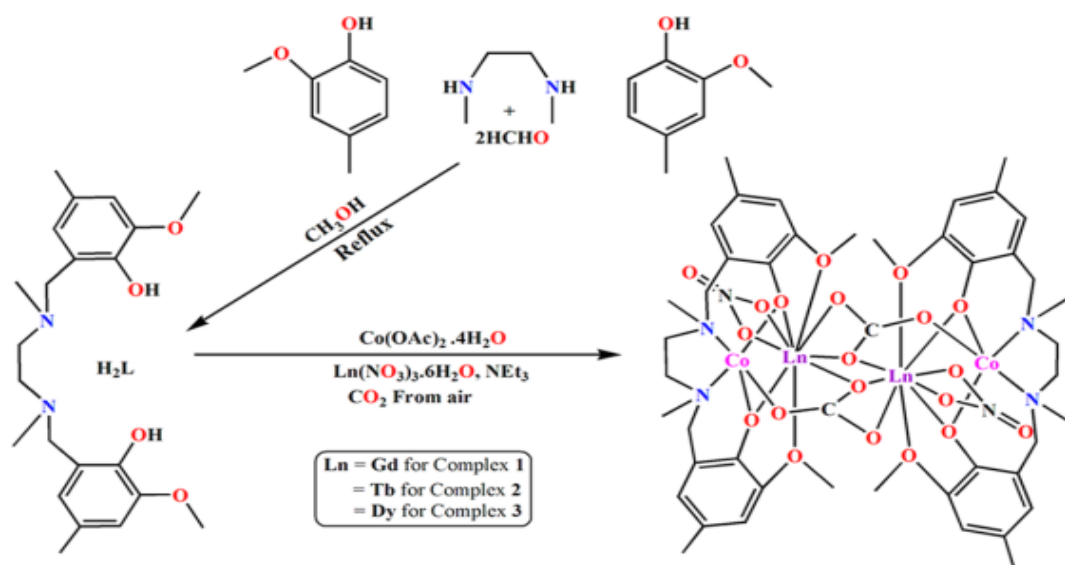


Figure 28: Syntheses of complexes 1–3 [58].

4. Conclusions

Through this article, Mannich-base ligands and their metal complexes were identified, with a study of their different applications in terms of their role in biological activity, as antioxidants, anticancer, urease, as catalysts for some reactions, and as corrosion inhibition for some metals in acidic, alkaline, or neutral mediums, as fluorescence sensors, and in various other fields such as magnetism and materials science.

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Conflict of Interest

The authors declare that they have no conflicts of interest.

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